**General Physical Science – Naming Compounds**

Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period\_\_\_\_\_\_\_\_\_\_

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| I can… |
| *Construct the name of binary ionic and covalent compounds using nomenclature rules.*  |

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| Bellwork |
| http://www1.whsd.net/courses/J0078/Periodic__Table/periodic_table.JPG | 1. What is the difference between an ionic bond and a covalent bond?\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_2. Give an example of an ionic bond.\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_3. Give an example of a covalent bond.\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| Notes |

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|  | 1. Chemists have a system for naming all types of compounds. Why would this be important?

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_1. Some compounds have common names that are used, like H2O is \_\_\_\_\_\_\_\_\_\_\_.
2. We will learn the steps for naming both ionic and covalent compounds. We will name only \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ compounds, meaning they have only ***two*** parts.
3. **These are the steps for naming binary *ionic compounds*:**
	* 1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ first, then \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**CaCl2****A. Calcium****B. Chlorine****C. Chlor + ide****Calcium chloride*** + 1. Write the metal name.
		2. Write the **root** of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and add \_\_\_\_\_\_\_\_\_\_\_ to the end.

Examples (write the names of the following):* + - 1. NaCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
			2. CaBr2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
			3. Na3P \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
1. **These are the steps naming binary *covalent compounds*:**
	1. The nonmetal farthest on the \_\_\_\_\_\_\_\_\_\_ of the table is written first.
	2. Add \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to indicate numbers of atoms.
	3. Never write the prefix “mono-“ for the \_\_\_\_\_\_\_\_\_\_\_ element.

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| --- | --- |
| ***Prefix*** | ***Number*** |
|  | 1 |
|  | 2 |
|  | 3 |
|  | 4 |
|  | 5 |
|  | 6 |
|  | 7 |
|  | 8 |
|  | 9 |
|  | 10 |

* 1. Change the ending to \_\_\_\_\_\_\_\_\_\_\_\_\_\_ for the second element.

**NO3****A. Nitrogen + oxygen****B. Mononitrogen trioxygen****C.** *~~Mono~~***Nitrogen trioxygen****D. Nitrogen trioxide(notice for oxygen, you drop the “y” and change to “i”!)**1. Try these covalent compounds on your own:

CCl4 = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_N2O = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_SF6 = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |