Name: _____

CHEMISTRY **SINGLE REPLACEMENT REACTION WORKSHEET**

Practice Reactions:

1. $Ag + KNO_3$ ----->

2. $Zn + AgNO_3$ ----->

3. $Al + H_2SO_4$ ----->

anion- use what you know about halogen reactivity...

4. $Cl_2 + KI$ ----->

5. $Li + H_2O$ ----->

6. $Cu + FeSO_4$ ----->

7. $Na + H_2O$ ----->

8. $Fe + Pb(NO_3)_2$ ----->

9. $Cu + H_2O$ ----->

10. $Cu + Al_2(SO_4)_3$ ---->

CHEMISTRY **SINGLE REPLACEMENT REACTION WORKSHEET**

11. Al + $Pb(NO_3)_2$ ---->

*12. Cl_2 + NaI --->

13. Fe + $AgC_2H_3O_2$ --->

14. Al + $CuCl_2$ --->

*15. Br_2 + CaI_2 --->

16. Al + HCl --->

17. Mg + HCl --->

18. Zn + H_2SO_4 --->

19. Fe + $CuSO_4$ --->

20. Cl_2 + MgI_2 --->

Accelerated chemistry

Name:_____

Solubility Rules Worksheet

Period:_____

1. Classify each of the substances as being soluble or insoluble in water.

   a. KBr =
   b. $PbCO_3$ =
   c. zinc hydroxide =
   d. sodium acetate =
   e. silver iodide =
   f. zinc carbonate =
   g. silver acetate =
   h. copper (II) sulfide =
   i. $Mg_3(PO_4)_2$ =
   j. KOH =
   k. NH_4OH =
   l. Hg_2SO_4 =
   m. PbI_2 =

2. Identify the two new compounds which form if the solutions, as suggested by the following table, were mixed. CIRCLE the names of the compounds which would precipitate from the solutions.

	KBr	Na_2CO_3	CaS	NH_4OH
$AgNO_3$				
$BaCl_2$				
$Al(NO_3)_3$				
$CuSO_4$				

3. a) ammonium chloride + silver nitrate →

b) iron (III) chloride + potassium phosphate →

c) lithium sulfide + lead (II) bromide →

d) sodium carbonate + calcium sulfate →